

through a mixture of the gas with three volumes of electrolytic hydrogen. The contraction produced in the latter case could be measured by cubic centimeters, and water used as an absorbent of the ammonia colored red litmus paper deep blue, besides giving a strong ammonia reaction with Nessler's reagent. With dilute hydrochloric acid as an absorbent there was obtained an abundant precipitate of ammonium platinic chloride. In a Geissler tube under a pressure of 10^{-3} mm and less the gas afforded the fluted spectrum of nitrogen with great brilliancy."

"Throughout the whole list of analyses in which nitrogen has been estimated the most striking feature is the apparent relation between it and the UO_2 . This is especially marked in the table of Norwegian uraninites recalculated, from which the rule might almost be formulated that, given either nitrogen or UO_2 the other can be found by simple calculation. The same ratio is not found in the Connecticut varieties, but if the determination of nitrogen in the Branchville mineral is to be depended on, the rule still holds that the higher the UO_2 the higher likewise is the nitrogen."

Compounds of Argon.—Berthelot finds that argon combines with the vapor of benzene under the influence of the silent electric discharge giving a yellow resinous odorous substance condensed on the surface of the two glass tubes between which the electric action is exerted. This substance, when heated, is decomposed, leaving a bulky carbonaceous residue. The volatile products of decomposition turn litmus blue, indicating that an alkaline substance is contained in them. (From a translation of Berthelot's article in the *Chem. News*, **71**, 151.)

ERRATA.—In Professor Venable's article in the February number the name of Victor Meyer occurs in several places. It should read Lothar Meyer.